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reactants is the Brønsted-Lowry acid, and which is the Brønsted-Lowry base? Explain your answer. $\text{H}_2\text{CO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HCO}_3^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$. The H_2CO_3 is the Brønsted-Lowry acid because it donates a proton to the H_2O . The H_2O is the Brønsted-Lowry base because it accepts the proton. b.

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Assessment Chapter Test B

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